

Chemical Equilibrium

Section 18.1 Equilibrium: A State of Dynamic Balance

In your textbook, read about chemical equilibrium.

Complete each statement.

- When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to _____.
- A reaction that can occur in both the forward and the reverse directions is called a(n) _____.
- _____ is a state in which the forward and reverse reactions balance each other because they take place at equal rates.
- At equilibrium, the concentrations of reactants and products are _____, but that does not mean that the amounts or concentrations are _____.
- Equilibrium is a state of _____, not one of _____.

In your textbook, read about equilibrium expressions and constants.

For each statement below, write *true* or *false*.

- _____ The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value.
- _____ The equation $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ is an example of a homogeneous equilibrium.
- _____ If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.
- _____ The value for K_{eq} is constant only at a specific volume.
- _____ If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.
- _____ A heterogeneous equilibrium means that reactants and products are present in more than one state.
- _____ The product of the forward chemical reaction is HI, for the equilibrium expression:

$$K_{eq} = \frac{[HI]^2}{[H_2][I_2]}$$